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Chapter 16:

Solutions 16.1

Properties of Solutions.

Solution Formation The compositions of the solvent and the solute determine whether a substance will dissolve.

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### Properties Of

The factors that determine how fast a substance dissolves are stirring (agitation) temperature the surface area of the dissolving particles

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Chapter 15&16

Solutions 16.1

Properties of Solutions

; Slide 6 ; Exothermic  
and Endothermic

Processes 17.1 Factors  
Affecting Solubility 1.

The type of solvent and  
solute A.Ionic solids  
and polar molecules

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### Properties Of

will dissolve in polar solvents but not in non polar ones. Ex. in the ionic solid  $\text{NaCl}(s)$ , ...

## Answers

### **Chapter 16 Solutions**

#### **16.1 Properties of Solutions - [PPTX ...**

Chapter 16 Properties of Solutions. STUDY.

PLAY. Solubility.

Amount of solute that dissolves in a given quantity of a solvent at a specified

temperature and

pressure to produce



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Properties Of

saturated solutions.

Solutions

Worksheet  
Answers  
Homogeneous mixture  
made of a solute (what  
got dissolved) and the  
solvent (what did the  
dissolving)

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Chapter 16 Properties  
Of Solutions Answers In  
all solutions, whether  
gaseous, liquid, or  
solid, the substance

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Answers

present in the greatest amount is the solvent, and the substance or substances present in lesser amounts are the solute (s). The solute does not have to be in the same physical state as the solvent, but

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Chapter 16 - Solutions -  
16.1 Properties of  
Solutions - Sample  
Problem 16.1 - Page  
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Solutions - 16.1  
Properties of**  
*Page 11/26*

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Properties Of

**Solutions ...**

Chapter 16 - Solutions.

Jennie L. Borders.

Section 16.1 -

Properties of Solutions.

Solutions are

homogeneous mixtures

that can be solids,

liquids, or gases.

Remember: The

solvent dissolves the

solute. Three factors

that determine the rate

at which a solute

dissolves are stirring

(agitation),

temperature, and the

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Properties Of  
Solutions

particle size of the  
solute.

Worksheet  
Answers  
**Chapter 16 -  
Solutions**

Chapter 16 Video  
Assignment Name

\_\_\_\_\_ 16.1 Properties of  
Solutions Chemistry  
and You: Watch the  
video on how we use  
solutions every day.

Name three ways  
solutions are used in  
everyday life \_\_\_\_\_ A.

**SOLUTION FORMATION**

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video on Solution  
Formation:?

$v = e^{-2E_{\text{oy}}/Y_{\text{amg}}}$  1,

What happens when  
particles dissolve in  
water.

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webquest..docx -**

**Chapter 16 Video ...**

1. Chapter 16

“Solutions” Pre-AP

Chemistry Charles Page

High School Stephen L.

Cotton 2. Section 16.1

Properties of Solutions

OBJECTIVES:

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### Properties Of

-Identify the factors that determine the rate at which a solute dissolves. 3. Section

16.1 Properties of Solutions OBJECTIVES:

-Identify the units usually used to express the solubility of a solute. 4.

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16.1 Properties of

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Chapter 16

Properties Of

Solutions 16.2

Concentrations of

Solutions 16.3

Colligative Properties

of Solutions 16.4

Calculations Involving

Colligative Properties.

16.1 Properties of

Solutions > 2 Parts

(components) of a

solution: Solute: Gets

dissolved Solvent:

Does the dissolving. H

2

**Chapter 16**

Chapter 16 Solutions.

Page 17/26

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Chapter 16

Properties Of

Mixtures - a review. ...

Solution: a mixture of  
two or more

substances that is  
identical throughout  
(homogeneous) can be  
physically separated.

composed of . ...

PowerPoint

Presentation - Chapter  
13 Properties of

Solutions Last modified  
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**Chapter 13**

*Page 18/26*

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Chapter 16

Properties Of

## **Properties of Solutions**

Chemistry - Chp 16 -

Solutions - Notes 1.

Chapter 16 "Solutions"

2. Section 16.1

Properties of

SolutionsOBJECTIVES: •

Identify the factors that determine the rate at which a solute

dissolves. • Identify the units usually used to express the solubility of a solute.

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Properties Of

## **Solutions - Notes**

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16.3 Colligative

Properties of Solutions

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of Solutions. Solution

Formation. Solutions

are homogeneous

mixtures that may be

solid, liquid, or

gaseous. The

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compositions of the solvent and the solute determine whether a substance will dissolve.

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Chapter 16 Solutions

167 SECTION 16.1

PROPERTIES OF

SOLUTIONS (pages

471-477) This section

identifies the factors

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### Properties Of

that affect the solubility of a substance and determine the rate at which a solute dissolves. Solution Formation (pages 471–472) Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2. 1.

## **05 Chem GRSW**

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### **Properties Of**

### **Solutions Answers**

Chapter 16: Colligative  
Properties of Solutions

45 16-4. The mole  
fraction of  $(\text{NH}_4)_2\text{SO}_4$   
 $(\text{aq})$  is given by  $x_{(\text{NH}_4)_2\text{SO}_4} = \frac{n_{(\text{NH}_4)_2\text{SO}_4}}{n_{(\text{NH}_4)_2\text{SO}_4} + n_{\text{H}_2\text{O}}}$ . Because  $(\text{NH}_4)_2\text{SO}_4(\text{aq})$  is a  
strong electrolyte, it  
dissociates completely  
into  $\text{NH}_4^+(\text{aq})$  and  
 $\text{SO}_4^{2-}(\text{aq})$  ions.

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Assume a one kilogram  
solution. The number  
of moles of ions in one

...

Answers

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