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## **Ph Review Problems Answers**

pH Review Problems 1)  
What is the molarity of a solution that has 450 grams of sodium chloride in 800 mL of water? 2) What is the molarity of a solution

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that contains 100 grams of iron (II) nitrate in 2.4 liters of water? 3) What is the pH of a solution that contains  $2.4 \times 10^{-5}$  moles of hydrobromic acid in 0.5 L of water?

## **pH Review Problems - nclark.net**

$\text{pH} = \text{pK}_a + \log$   
(conjugate base/ acid)

$\text{pH} = 4.7 + \log$   
(0.1/0.2) = 4.7 - 0.3.

$\text{pH} = 4.4$ . 3. For a weak acid with a  $\text{pK}_a$  of 6.0,

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show how you would calculate the ratio of acid to salt at pH 5.

Ans: 4. Suppose you have just added 100 mL of a solution containing 0.5 mol of acetic acid per liter to 400 mL of 0.5 M NaOH.

## **pH Practice Problems with Answers ~ Biology Exams 4 U**

Solution:  $\text{pH} = -\log$   
 $[\text{H}^+] = -\log (5.31 \times 10^{-9}) = 8.27$ . Example

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3: Calculate  $[H^+]$  for a solution having a pH of 1.57. Solution:  $[H^+] = 10^{-pH} = 10^{-1.57} = 0.0269 \text{ M}$ , or  $[H^+] = \text{antilog}(-pH) = \text{antilog}(-1.57) = 2.69 \times 10^{-2} \text{ M}$ . To perform the antilog function on most calculators, use  $10^x$  or  $10^{-x}$ .

## **pH Problems - VCC Library**

Problem : What is the pH of a 0.001 M solution of  $H_2SO_4$

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$\text{HSO}_4^-$  has a  $\text{pK}_a$  of  $1.2 \times 10^{-2}$ . To solve this problem, you must first note that sulfuric acid's first deprotonation is as a strong acid, so we have a concentration of  $0.001 \text{ M H}^+$  to start and  $0.001 \text{ M}$  hydrogen sulfate. Because hydrogen sulfate is a weak acid, this problem becomes very similar to the last one (see ).



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## **pH Calculations: Problems and Solutions | SparkNotes**

pH Problems

Worksheet: Answers 1.

e By definition, a solution with a pH less than 7 is an acid and has a higher concentration of  $H^+$  than  $OH^-$ . The closer the pH gets to 0, the more acidic it is; thus a solution with  $pH = 2$  is highly acidic. 2. d  $pH = -\log_{10} [H^+]$ .

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## **pH Problem Worksheet Answers - pH Problems Worksheet ...**

Solutions to Review Problems for Acid/Base Chemistry 3. If 13.2 g  $\text{NaC}_2\text{H}_3\text{O}_2$  (FW = 82.0) are added to the 800 mL of solution in Problem 2, what is the resulting pH? The addition of  $\text{C}_2\text{H}_3\text{O}_2^-$  to a solution of  $\text{HC}_2\text{H}_3\text{O}_2$  creates a  $\text{HC}_2\text{H}_3\text{O}_2 / \text{C}_2\text{H}_3\text{O}_2^-$

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2-buffer. initially,  $[H_2CO_3] = 0.195\text{ M}$   
and  $\text{mol } C_2H_3O_2^- =$   
 $13.2\text{ g}/82.0\text{ g/mol} =$   
 $0.161\text{ mol } [C_2H_3O_2^-]$

## **Solutions to Review Problems for Acid/Base Chemistry**

Test your knowledge  
on pH, acids, and  
bases! If you're seeing  
this message, it means  
we're having trouble  
loading external  
resources on our  
website. ... pH, acids,

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and bases review.

Practice: pH, acids, and bases. This is the currently selected item. pH, acids, and bases review. Biology is brought to you with support from the Amgen Foundation.

## **pH, acids, and bases (practice) | Khan Academy**

pH practice - Answers.

1) What is the pH and pOH of a  $1.2 \times 10^{-3}$  HBr solution? pH: 2.9

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pOH: 11.1. 2) What is the pH and pOH of a  $2.34 \times 10^{-5}$  NaOH solution? pOH: 4.6 pH: 9.4. 3) What is the pH and pOH of a solution made by adding water to 15 grams of hydroiodic acid until the volume of the solution is 2500 mL? pH: 1.6 pOH: 12.4

## **Acid and Base Worksheet - Answers**

[Return to Question.](#)

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Interpret the following  
ABG Values to  
determine what type of  
Acid-Base Imbalance is  
present.

Uncompensated  
examples. 1) pH: 7.30,  
PaCO<sub>2</sub>: 38, HCO<sub>3</sub><sup>-</sup>:  
18 = Metabolic  
Acidosis 2) pH: 7.25;  
PaCO<sub>2</sub>: 50; HCO<sub>3</sub><sup>-</sup>:  
23 = Respiratory  
Acidosis 3) pH: 7.49;  
PaCO<sub>2</sub>: 33; HCO<sub>3</sub><sup>-</sup>:  
25 = Respiratory  
Alkalosis Partially  
Compensated

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examples

## **ABG Practice Answers - Part 1 - Your Nursing Tutor**

Answer:  $\text{pH} = -\log(0.0001) = 4$ . Usually, you aren't given the hydrogen ion concentration in a problem but have to find it from a chemical reaction or acid concentration. The simplicity of this will depend on whether you have a strong

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acid or a weak acid.

## **Here's How to Calculate pH Values - ThoughtCo**

$\text{pH} = -\log(2 \times 10^{-5}) = 4.7$ . These problems reduce to a very simple form since the value of X depends on  $K_a$  and the initial ratio of A-/HA. Thus, unlike the other two classes of problems, the value of X does not depend on the actual concentrations of A-



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and HA (provided both A-and HA are large enough that the X can be ignored).

## **ACID-BASE BUFFER PROBLEMS**

pH value less than 7  
Bases Taste bitter Feel slippery Contain a hydroxide ion ( $\text{OH}^-$ )  
pH value greater than 7 BOTH change colors of indicators react with each other to form salt and water conduct electricity when

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dissolved in solution  
(electrolytes) 2.

## **Exam #10 Review: Acids, Bases, and pH**

Be ready to answer how long the problem has been going on, as well as specifics such as the temperature of the solution. Of course, if the pH tester has any cracks or physical damage, this likely is diminishing its performance. At this point, the technical

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expert can propose the next best steps. Last Updated: 03/05/20

## **pH Meter Troubleshooting Procedures & Calibration - Cole ...**

Now we can find the pOH. The sum of the pH and the pOH is always 14. The pOH of the solution is 7.8.

Alternatively, a shortcut can be used to estimate the pH. If is in the form , then pH is

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roughly . For this question, this shortcut gets us a pH of 6.4, which produces a pOH of 7.6; very close to the real answer!

## **Calculating pH and pOH - High School Chemistry**

Ksp Problems -

Chemistry Name: \_\_\_\_\_

1) The value of Ksp of AgCl is  $1.8 \times 10^{-10}$ .

What would be the molar concentration ...

Adjust pH at 11.91

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Ca<sup>+2</sup> does not precipitate but Mg<sup>+2</sup> will ppt Sample Problem #9 Determine optimum conditions to separate 0.10M Ni<sup>+2</sup> & 0.10M Sr<sup>+2</sup> by precipitating with Na<sub>2</sub>CO<sub>3</sub>

## **Ksp Problems - Chemistry**

Question: Problem 07  
(Review Problems Of  
Water Chemistry  
Related To  
Atmosphere/Climate) It

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Has Been Estimated That The Concentration Of CO<sub>2</sub> In The Atmosphere Before The Industrial Revolution Was About 275 Ppm. If The Accumulation Of CO<sub>2</sub> In The Atmosphere Continues, Then The Middle Of This Century, It Will Probably Be Around 600 Ppm. Calculate The PH Of Rainwater ...

**Solved: Problem 07  
(Review Problems Of**

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Answers  
**Water Chemistry Rel**

...

30. The solution is alkaline with  $\text{pH} = 8.34$ . 31. The solution required 0.056 mole of acetic acid. From the  $\text{pH}$ ,  $[\text{H}^+] = 10^{-3}$  and  $[\text{CH}_3\text{COO}^-]$  must be the same. 32. The conjugate base of is the carbonate ion, formed by the loss of a proton. The conjugate acid is carbonic acid  $\text{H}_2\text{CO}_3$ , formed as gains a proton. 33. 34.

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## **Answers to Chemistry Problems**

All that remains that affects the pH at the equivalence point is the conjugate base of the weak acid,  $C_2H_3O_2^-$ . This is a weak base equilibrium problem because the conjugate bases of all weak acids are weak bases themselves.  $C_2H_3O_2^- + H_2O \rightleftharpoons HC_2H_3O_2 + OH^-$



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**AP Chemistry**  
**Chapter 15 Answers**  
**- Zumdahl 15**

This chemistry video tutorial explains how to calculate the pH of a buffer solution using the henderson hasselbalch equation. It explains the concept, compon...

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*Page 25/26*

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